1. Balanced Equation: 2KClO₃ → 2KCl + 3O₂

What is the mass of A given?*15 g KClO₃*

Which element/compound(B) are you solving for? *O2*

|  |  |  |  |
| --- | --- | --- | --- |
| *15 g KClO₃* | *1 mol KClO₃* | *3 mol O2* | *32 g O2* |
|  | *122.55g KClO₃* | *2 mol KClO₃* | *1 mol O2* |

Answer: *5.88 g O2*

2. Balanced Equation: 4NH₃ + 5O₂ → 4NO + 6H₂O

What is the mass of A given? *30 g NH₃*

Which element/compound(B) are you solving for? *NO*

|  |  |  |  |
| --- | --- | --- | --- |
| *30 g NH₃* | *1 mol NH₃* | *4 mol NO* | *30.01 g NO* |
|  | *17.04 g NH₃* | *4 mol NH₃* | *1 mol NO* |

Answer: *52.83 g NO*

3. Balanced Equation: 2Al + 3Br₂ → 2AlBr₃

What is the mass of A given? *50 g Al*

Which element/compound(B) are you solving for? *AlBr₃*

|  |  |  |  |
| --- | --- | --- | --- |
| *50 g Al* | *1 mol Al* | *2 mol AlBr₃* | *266.78 g AlBr₃* |
|  | *26.98 g Al* | *2 mol Al* | *1 mol AlBr₃* |

Answer: *494.40 g AlBr₃*