EXIT TICKET

**Suppose you're in a chemistry lab, and you want to synthesize water (H2O) by reacting hydrogen gas (H2) with oxygen gas (O2) according to the following balanced chemical equation:**

**2H2(g) + O2(g) -> 2H2O(g)**

**You have 10.0 grams of hydrogen gas (H2) and 40.0 grams of oxygen gas (O2). After carrying out the reaction, you find that you obtain 16.0 grams of water (H2O). Calculate the percent yield of the reaction.**

|  |
| --- |
|  |
|

----------------------------------------------------------------------------------------------------------------------------

EXIT TICKET

**Suppose you're in a chemistry lab, and you want to synthesize water (H2O) by reacting hydrogen gas (H2) with oxygen gas (O2) according to the following balanced chemical equation:**

**2H2(g) + O2(g) -> 2H2O(g)**

**You have 10.0 grams of hydrogen gas (H2) and 40.0 grams of oxygen gas (O2). After carrying out the reaction, you find that you obtain 16.0 grams of water (H2O). Calculate the percent yield of the reaction.**

|  |
| --- |
|  |
|